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what ' s in matter. The emission of light by hydrogen and other atoms has played a key role in understanding the electronic structure of atoms. Trace materials, such as evidence from a crime scene, lead in paint or mercury in drinking water, can be

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identified by heating or burning the materials and examining the color (s) of light given off in the form of bright-line spectra.

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Electron Energy and Light 3 11.

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Does a gas discharge tube filled with boron emit the same wavelengths of light as a tube filled with hydrogen? Use evidence from Model 2 to support your answer. 13.

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Which color corresponds to the shortest wavelengths? violets c. Write a sentence that describes the relationship between wavelength and energy of light. the wavelength gets shorter and light energy increases Electron Energy and Light 1 Model 2 – Emission

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Spectra for Hydrogen and Boron
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Electron energy and light answers pdf When an electron is struck by a photon of light, it absorbs the quantum of energy that the photon carries and moves to a higher energy state. One way of thinking

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about this higher energy state is to imagine that the electron is now moving faster (it has just been hitting a fast-moving photon).

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Light Energy. Each orbital has a

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specific energy associated with it. For an electron to be boosted to an orbital with a higher energy, it must overcome the difference in energy between the orbital it is in, and the orbital to which it is going. This means that it must absorb a photon that contains precisely that

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amount of energy, or take exactly that amount of energy from another particle in a collision.

~~Background: Atoms and Light Energy – NASA~~

The electron is a subatomic particle, symbol e^- or e^{-} ,

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whose electric charge is negative one elementary charge. Electrons belong to the first generation of the lepton particle family, and are generally thought to be elementary particles because they have no known components or substructure. The electron has a

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mass that is approximately $1/1836$ that of the proton.

~~Electron~~ — Wikipedia

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The energy level of the electron of a hydrogen atom is given by the following formula, where.

$E_n = -\frac{1312}{n^2} \text{ kJ/mol}$. E_n . kJ/mol. For a single electron instead of per mole, the formula in eV (electron volts) is

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also widely used: $E_n = -13.6 \frac{eV}{n^2}$. E_n .

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The photoelectric effect is the emission of electrons when

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electromagnetic radiation, such as light, hits a material. Electrons emitted in this manner are called photoelectrons. The phenomenon is studied in condensed matter physics, and solid state and quantum chemistry to draw inferences about the properties of

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atoms, molecules and solids. The effect has found use in electronic devices specialized for light detection and precisely timed electron emission. The experimental results disagree with

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