

Chapter 11 Chemistry Answers

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11. molar mass = mass/ moles; mass = 45.0 g; find moles; $T_b = 90.2^\circ\text{C} - 80.1^\circ\text{C} = 10.1^\circ\text{C}$ $\left(\text{moles} = \frac{\Delta T_f}{2.53} \times \text{kg}\right)$:benzene = $\frac{10.1 \times 10^{-3}}{2.53} \times 0.0750 = 0.300 \text{ moles}$

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Answer/Explanation. Answer: Explanation: (i) (d) (ii) (e) (iii) (b) (iv) (a) In cleavage of ethers, higher alcohol and lower alkyl halide is formed except in tert. butyl methyl ether, in which (CH₃)₃C+ is more stable so (CH₃)₃C-I and CH₃OH is formed. Anisole with HI gives phenol and CH₃I because phenoxide ion is stabilised by resonance.

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