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11. Ionization Energy Ionization Energy: (of an atom) is the energy required to remove an electron from an individual atom in the gas phase -As we have noted, the most characteristic chemical property of a metal atom is losing electrons to nonmetals. -Another way of saying this is to say that metals have relatively low ionization energies;

~~Chapter 11: Modern Atomic Theory~~
~~Chemistry SA Akpontier~~

414 Chapter 11 Modern Atomic Theory
11.1 The Mysterious Electron 415 When a guitar string is plucked, the string vibrates up and down in a wave pattern.

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Review Skills 11.1 The Mysterious
Electron Standing Waves and Guitar
Strings Electrons as Standing Waves

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Waveforms for Hydrogen Atoms Particle Interpretation of the Wave Character of the Electron Other Important Waveforms Overall Organization of Principal Energy Levels, Sublevels, and Orbitals 11.2 Multi-Electron Atoms.

~~Chapter 11 Modern Atomic Theory—An Introduction to Chemistry~~

Chapter 11. Light travels as a photon (particle) and/or a wave. Energy: Atoms become excited (excited state), emit photons, crash (ground state), and produce color. Rutherford's atom: The nucleus consists of protons and electrons, the atom is mostly empty, and the positive core is surrounded by negative electrons.

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Chapter 9; Chapter 11-Modern Atomic Theory; Chapter 12-Chemical Bonding;

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Chapter 15; Chapter 14 and 16; Major Points *Energy travels through space by electromagnetic radiation ("light"), which can be characterized by the wavelength and frequency of the waves. Light can also be thought of as packets of energy called photons.

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Chapter 11: Modern Atomic Theory. Main Concept: Electromagnetic Radiation . Energy is being transmitted from one place to another by light ; Different types ; Wavelength - the distance between two consecutive wave peaks ; Frequency - indicates how many wave peaks pass a certain point per given time period ;

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Chapter 12: Chemical Bonding; Chapter 13/15: Gases and Solutions; Chapter 14: Liquids and Solids; Chapter 16: Acid/Base Chemistry; Terms: Electromagnetic Radiation- A form of energy that exhibits wave-like behavior as it travels through space and time

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Scientist believed atoms can exist on any energy levels because only certain photons (particles of light) are emitted, We know that only certain energy changes are occurring and the level the electrons are on. There are certain discrete energy level that will always emit photons (Quantized vs. Continuous) Rutherford's Atom- Electrons orbit nucleus and nucleus is made up of protons and neutrons.

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STUDY. PLAY. Bohr. he is known as the father of Quantum Physics. He equated the wavelength of an emitted photon with distance. He believed in a planetary model of an atom and his model asserts that:

Louis de broglie.

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Modern Chemistry Chapter 11 416

Chapter 11 Modern Atomic Theory

objeCtive 3. 11.1 The Mysterious Electron

417 the electron are also called orbitals.

The orbital shown in Figure 11.3 is called the 1s orbital. The negative charge is most intense at the nucleus and decreases in intensity

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Chapter 11: modern atomic theory. wave, light, energy levels, bohr model of atoms, wave mechanical model, s orbitals and d orbitals, Pauli Exclusion Principle, Ionization Energy. Review Rutherford's atom model states:-electrons orbit a positive core made up of neutrons and protons

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~~MK Schuster SAA Chemistry~~

Chapter 11-Modern Atomic Theory.
Chapter 12-Chemical Bonding. Chapter 5-Nomenclature. Chapter 6/7-Reactions. Chapter 8-Chemical Composition Chapter 9-Chemical Quantities. Chapter 10-Energy. Chapter 13-Gas and Chapter 15-Solutions. Chapter 16-Acid Base and Chapter 14-Phases. Review. In Rutherford's atom electrons orbits a small positive core ...

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Discovery of the Electron. Discovery of the Nucleus. Nova Science Now - Atomic Structure Short. Ted Ed - The Uncertain Location of Electrons. Rutherford Experiment. ChemMatters - Fireworks!

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Chapter 11: Modern Atomic Theory ;
Chapter 13: Gases; Chapter 14: Liquids

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and Solids ; Chapter 15: Solutions;

Chapter 16: Acids and Bases; About.

Electromagnetic radiation: a term used to describe all the different kinds of energies released by the Sun. For more help with this concept visit this website _

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~~J.G.M.C.K.~~

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1 Chapter 11 Modern Atomic Theory 2

Electromagnetic Radiation. Light is a form of electromagnetic (EM) radiation ; All forms of EM radiation are types of kinetic energy ; See page 306; 3 EM Radiation.

EM radiation can be described as traveling in waves or as packets of energy called photons; 4 EM Radiation

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The modern atomic theory states that

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atoms of one element are the same, while atoms of different elements are different. What makes atoms of different elements different? The fundamental characteristic that all atoms of the same element share is the number of protons. All atoms of hydrogen have one and only one proton in the nucleus; all atoms of iron have 26 protons in the nucleus.

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Chemistry Traveling with the Atom

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