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Guide Answers

States of Matter (chapter 13)

Flashcards | Quizlet

You are already familiar with the three common states of matter: solid, liquid, and gas. Solid objects litter the room around you. For example, you can easily

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recognize the shape of your desk; you know that your backpack cannot hold seven textbooks. You encounter liquids throughout the day as yo u

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SECTION 13.1 THE NATURE OF GASES (pages 385–389) This section introduces the kinetic theory and describes how it applies to gases. It defines gas pressure and explains how temperature is related to the kinetic energy of the particles of a

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substance. Kinetic Theory and a
Model for Gases (pages 385–386)
1.

*Name Date Class STATES OF
MATTER 13*

There are three states of matter
that we will learn about in this

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chapter. (If you want to learn about more states of matter, I can refer you to somebody.) Those three states are solid, liquid, and gas. These three states are quite different. The main difference is in their particles.

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applies to gases. It defines gas pressure and explains how temperature is related to the kinetic energy of the particles of a substance.

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SECTION 13.1 THE NATURE OF
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Section introduces the kinetic theory and describes how it applies to gases. It defines gas pressure and explains how temperature is related to the kinetic energy of the particles of a substance.

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Chapter 13 States Of Matter Worksheet

Chapter 13: States of Matter.
STUDY. PLAY. Kinetic Molecular
Theory. Explains the properties of
gases in terms of the energy,
size, and motion of their particles.
Elastic Collision. Describes a

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collision in which kinetic energy may be transferred between colliding particles but the total kinetic energy of the two particles remains the same.

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Flashcards | Quizlet*

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13.1 The Nature of Gases - 13.1
Lesson Check - Page 424 8
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0132525763, ISBN-13:

978-0-13252-576-3, Publisher:
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*Chemistry (12th Edition) Chapter
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all matter consists of tiny
particles that are constantly in

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motion What are the three assumptions of the kinetic theory as it applies to gases? -The particles in a gas are considered to be small, hard spheres with an insignificant volume. -The motion of the particles in a gas are rapid, constant, and random.

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recognize that ending poverty ...

*United Nations Sustainable
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13.1 The Nature of Gases -

Chemistry & You; 13.1 The Nature
of Gases - Sample Problem 13.1;

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13.1 The Nature of Gases - 13.1
Lesson Check; 13.2 The Nature of
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The Nature of Liquids - 13.2
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Solids - Chemistry & You; 13.3
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Check

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Title: Chapter 13 States of Matter
1 Chapter 13 States of Matter 2
Kinetic Theory as Applied to
Gases Fundamental assumptions
about gases. The particles in a

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gas are considered to be small, hard spheres with an insignificant volume. Between particles in a gas there is empty space. No attractive or repulsive forces exist between the particles. 3

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Worksheet

The attacks the USPS continues to face are not just attacks on the postal service but attacks on Black lives. To defund the USPS would be to deny future generations this opportunity and dishonor the legacy of Black

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postal workers. Now, we're taking this matter into our own hands by writing and sending #BlackLoveLetters through USPS...

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