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where k is the rate constant and n is the order of the reaction.; Half life of a reaction, $t_{1/2}$, is the time it takes for one-half of a quantity of a reactant to react. According to the question the

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time taken for the concentration of the reactant to decrease from $[A]_0$ to $[A]_0/2$ is the same as the time taken to decrease the concentration from $[A]_0/2$ to $[A]_0/4$.

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Ans. $\frac{1}{3} \text{ m} \mid \frac{2}{3} \text{ m} = r \text{ A} \mid = \frac{1}{3} r \text{ A} \mid \frac{3}{3} = \text{L} \mid 0 \times 2 \text{ (r A dx)}$

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