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~~PV Diagrams, Internal Energy, Heat, Work, Isothermal, Adiabatic, Isobaric, Physics AP Chemistry Unit 5 Part 1 Review: Reaction Kinetics Calorimetry Hatchet Chapter 17 Calorimetry: Crash Course Chemistry #19 Gibbs Free Energy - Equilibrium Constant, Enthalpy \u0026 Entropy - Equations \u0026 Practice Problems Chemical Kinetics Rate Laws – Chemistry Review – Order of Reaction \u0026 Equations Enthalpy Change of Reaction \u0026 Formation - Thermochemistry \u0026 Calorimetry Practice Problems Enthalpy: Crash Course Chemistry #18 AP Chemistry Unit 4 Review: Chemical Reactions Standard Enthalpy Of Formation - Thermodynamics (Part 17)~~

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You must show all work and setup for this to count as extra credit on your test (+3 points) 15.1 – 15.2 Heat, Calorimetry, and Enthalpy.

Make the following conversions:  
444 cal to joules. 1.8 kJ to joules.  
0.45 kJ to calories.

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### THERMOCHEMISTRY SECTION

#### 17.1 THE FLOW OF ENERGY-

#### HEAT AND WORK (pages

505-510) This section explains

the relationship between energy

and heat, and distinguishes

between heat capacity and specific

heat. ~ Energy Transformations

(page 505) 1. What area of study

in chemistry is concerned with the

heat transfers that occur during

chemical reactions?

thermochemistry

### ~~THERMOCHEMISTRY~~

Chapter 17 Thermochemistry 183

#### SECTION 17.1 THE FLOW OF

#### ENERGY—HEAT AND WORK

(pages 505 – 510) This section

explains the relationship between

energy and heat, and distinguishes

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between heat capacity and specific heat. Energy Transformations (page 505) 1. What area of study in chemistry is concerned with the heat transfers that

~~SECTION 17.1 THE FLOW OF ENERGY HEAT AND WORK (pages 505–510)~~

Chapter 17 Thermochemistry Study Guide - Weebly Study Guide Thermochemistry  $Q = m C \Delta T$

$H = -q C = 4184 \text{ J moles goC } 1$   
When 150-g sample of KCl dissolves in 650 g of water in a calorimeter, the temperature drops from  $310^\circ \text{C}$  to  $205^\circ \text{C}$

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this chapter we will review some  
of Chapter 6, Thermochemistry,  
and study a part of Chapter 6  
which we omitted in General  
Chemistry I. Everything in the  
universe is energy or matter.

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