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1. In covalent bonds, electron sharing usually occurs so that atoms attain the electron configuration of noble gases. (8) 2. Atoms form double or triple covalent bonds if they can attain a noble gas structure by sharing two pairs or three pairs of

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electrons.

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Chapter 8 Covalent Bonding and Molecular Structure 8-3 There are two types of repulsive forces between the two atoms. First, the nuclei repel because they are both positively charged. Second, the electrons repel because they are both negatively charged. The attractive forces between the two atoms result from the

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## **Structure**

1) Ionic bonds—transfer of electrons. (gained or lost; makes formula unit) 2) Covalent bonds—sharing of electrons. The resulting particle is called a “molecule”. Covalent Bonds. • The word covalent is a combination of the prefix co- (from Latin com, meaning “with” or “together”), and the verb. valere, meaning “to be strong”.

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Bonding Covalent Bonding G. Lewis (1916) •  
Some atoms share e<sup>-</sup> to form

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A covalent bond in which the bonding electrons are most likely to be found in sausage-shaped regions above and below the bond axis of the bonded atoms. Polar molecule. A molecule in which one side of the molecule is slightly negative and the opposite side is slightly positive. Polar covalent bond.

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Bonding Covalent Bonding Solutions Manual  
Chemistry: Matter and Change • Chapter 8 121  
Section 8.1 The Covalent Bond pages 240–247  
Practice Problems page 244 Draw the Lewis  
structure for each molecule. 1. PH<sub>3</sub> H<sub>2</sub> H<sub>2</sub> H<sub>2</sub>  
H<sub>2</sub> P<sub>2</sub> respectively, for single, double, and  
triple P – – 2. H<sub>2</sub> S H<sub>2</sub> H<sub>2</sub> H<sub>2</sub> – H<sub>2</sub> S S – 3. HCl

## Covalent Bonding Covalent Bonding – Weebly

Complete Table 8.10 , which shows the number  
of electrons shared in a single covalent

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bond, a double covalent bond, and a triple covalent bond. Identify the group of atoms that will form each of these bonds.

## **Covalent Bonding | Glencoe Chemistry: Matter and**

View full document. Chapter 8 Covalent Bonding. The Covalent Bond Atoms will share electrons in order to form a stable octet. Covalent bond : the chemical bond that results from the sharing of valence electrons also called a molecular bond. Molecule : formed when two or more atoms bond covalently Formed when the proton of one atom is

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attracted to the electron cloud of another atom.

## **6613055.ppt - Chapter 8 Covalent Bonding The Covalent Bond ...**

242 Chapter 8 • Covalent Bonding Single Covalent Bonds When only one pair of electrons is shared, such as in a hydrogen molecule, it is a single covalent bond. The shared electron pair is often referred to as the bonding pair. For a hydrogen molecule, shown in Figure 8.4, each covalently bonded atom equally attracts the pair of shared electrons.

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Chapter 8 2. 

- The atoms held together by sharing electrons are joined by a Covalent

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Bond . </li></ul>Sharing is Caring 3. 2.  
Covalent bonds- Two atoms share one or more  
pairs of outer-shell electrons. Oxygen Atom  
Oxygen Atom Oxygen Molecule (O<sub>2</sub>) 4.

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