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Monatomic Ion. Consists of a single atom with a positive or negative charge resulting from the loss of gain of one or more valence electrons and behaves as a unit, yet it still has a charge. (+ or -). Determined by the amount of electrons lost or gained.

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29 terms. hales913. Chapter 9:
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ion. - single atom with a positive
or negative charge resulting from
the loss or gain of electrons,
respectively.

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Section 9.4 – Naming and Writing
Formulas for Acids and Bases. An
acid is a compound that produces
 H^+ ions when it dissolves in
water. The formula for an acid
normally starts with an H. When
naming acids, you should first
determine whether or not the

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Chapter 9 Chemistry Chemical
Names and Formulas. Acids.

base. Cation. Ionic compounds.

Compounds that contain one or
more hydrogen atoms and

produce.... An ionic compound

that produces hydroxide ions

when dissolved.... Any atom or

group of atoms that has a positive

charge. Compounds composed of

cations and anions.

~~chemistry vocabulary chapter 9~~

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~~Chapter 9 Chemical Names and
Formulas~~ 83 SECTION 9.3 NAMING
AND WRITING FORMULAS FOR
MOLECULAR COMPOUNDS (pages
268–270) This section explains
the rules for naming and writing
formulas for binary molecular
compounds. Naming Binary
Molecular Compounds (pages
268–269) 1. Circle the letter of
the type(s) of elements that form
binary molecular compounds.

~~Name Date Class~~ CHEMICAL
NAMES AND FORMULAS 9
SECTION 9.2 NAMING AND
WRITING FORMULAS FOR IONIC
COMPOUNDS I. Write the formulas
for these binary ionic compounds.
c. potassium iodide 2. write the
formulas for the compounds

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formed c. b. f. 3. Name the following binary ionic compounds.
sodium sulfide g. CuCl : h. SnCl_2 i. a.
 MnO_2 c. d. SrBr_2 221 Chapter 9
Chemical Names and Formulas

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Chapter 9 Chemical Names and Formulas
9.1 Naming Ions
9.2 Naming and Writing Formulas for Ionic Compounds
9.3 Naming and Writing Formulas for Molecular Compounds
9.4 Naming and Writing Formulas for Acids and Bases
9.5 The Laws Governing How Compounds Form

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Chapter 9 "Chemical Names and Formulas" ... in parenthesis after

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The name of the metal (Table 9.2, p.255) Predicting Ionic Charges
Some of the post-transition elements also have more than one possible oxidation state. Tin (II) = Sn^{2+} Lead (II) = Pb^{2+} Tin (IV) = Sn^{4+} Lead (IV) = Pb^{4+}

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281 CHAPTER 9 Assessment 42.

a. 2- b. 1+ c. 1- d. 3+ 43. a. 2+

b. 2+ c. 3+ d. 1+ 44. a. barium
ion b. iodide ion c. silver ion d.

mercury(II) ion 45. cyanide, CN^-
and hydroxide, OH^- 46. a.

hydroxide ion b. lead(IV) ion c.

sulfate ion d. oxide ion 47. zero; A
compound is electrically neutral.

48. The symbols for the cation
and

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Chemical Names and Formulas.

Section 9.1: Naming Ions. 1. What

is the charge on the ion typically
formed by each element? a.

oxygen c. sodium e. nickel, two
electrons lost . b. iodine d.

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aluminum f. magnesium. 2. How many electrons does the neutral atom gain or lose when each ion forms? a. Cr^{3+} c. Li^+ e. Cl^- b. P^{3-}

...

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Formulas" Tools. Copy this to my account; E-mail to a friend; Find other activities; ... In samples of any chemical compound, the masses of the elements are always in the same proportions any atom or group of atoms that has a negative charge ... place cation name first followed by the anion name: naming polyatomic ...

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