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*METAL In Your HAND! - Liquid*  
*Metal Science Experiments 9*  
*Awesome Science Tricks Using*  
*Static Electricity! Amazing*  
*Balloon Tricks and Science*  
*Experiments* ~~How to make a~~

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working model of a wind  
turbine from cardboard |  
school project Electrolysis  
of water experiment using  
pencils, h2o electrolysis,  
electrolysis water~~

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Steam Power Generator -  
Science Project ~~Newton's~~  
~~First Law of Motion~~ Class  
~~9 Tutorial Class 9 Science~~  
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## 56 Answers

ChemActivity 56 The  
Equilibrium Constant (III)  
3275. Assume that a liquid  
boils at the temperature at  
which  $G^\circ = 0$  for the  
reaction liquid gas. a) Adding  
salt to water does not  
change  $H^\circ$  for the process

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$H_2O(l) \rightleftharpoons H_2O(g)$  However,  $S^\circ$   
for the process is decreased  
because the entropy of the  
liquid is

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Equilibrium Constant (III)  
3275. Assume that a liquid  
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H<sub>2</sub>O(l) H<sub>2</sub>O(g) However,  $S^\circ$   
for the process is decreased  
because the entropy of the  
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The Equilibrium Constant  
(III) 3275. Assume that a  
liquid boils at the  
temperature at which  $G^\circ = 0$   
for the reaction  
liquid  
gas. a) Adding salt to water  
does not change  $H^\circ$  for the  
process  $H_2O(l) \rightleftharpoons H_2O(g)$

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ChemActivity 19 1.a) b) four  
c) tetrahedral d) about 109  
e) bent f)  $sp^3$  2. carbonate,  
see Ex 4 in CA 17, trigonal  
planar,  $sp^2$

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answers, it is your  
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office hours or ask during  
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HW Practice Keys -  
ChemActivity 51: Cell  
Voltage - ChemActivity 50:

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- HW Practice Keys -  
ChemQuest 55: Free Energy -  
ChemQuest 54: 2nd Law of  
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College ChemActivity 4 1.  
No. The ionization energy of  
He would be about  $4\times$  (twice  
the charge and half the  
distance) the ionization  
energy of H if this were the

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case. 2. Several possible answers. A typical one would have all three electrons at a much farther distance than in H from the

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Energy and Equilibrium •  
Know that the Gibbs energy was developed as a single energy term that incorporates both the enthalpy and entropy effects on reaction spontaneity.  $\Delta G^\circ = \Delta H^\circ - T\Delta S^\circ$  • Be able to

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calculate  $\Delta G^\circ_{\text{rxn}}$  from the above equation (make sure your units match! kJ  $\neq$  J)

Chemistry General, Organic,  
and Biological Chemistry

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Edition (Hard Cover) Grade  
11 Chemistry 2e Introductory  
Chemistry Encyclopedia of  
Physical Organic Chemistry,  
6 Volume Set Specific Ion  
Effects Physical Chemistry,  
4th Edition Infochemistry



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Organic Chemistry Plasma  
Catalysis The European World  
1500–1800 The Shadow Prince  
Science Reporter The Carbon  
Cycle Free Energy Methods in  
Drug Discovery Physical  
Chemistry for the  
Biosciences Surface-Enhanced

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Raman Scattering Picture-  
Perfect Science Lessons  
Thermoplasmonics

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