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$\text{l}/1.25718 \text{ g}) = 0.16218 \text{ l} =$

$162.18 \text{ ml} = 162.18 \text{ cm}^3$ 2.

$(0.2096)(1.42952 \text{ g/L}) +$

$(0.7811)(1.25092 \text{ g/L}) +$

$(0.00930)X = 1.000(1.29327 \text{ g/L})$

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In addition, a recent tenure as a mentor of the U.S. Chemistry Olympiad Team, Professor Kotz has received numerous honors, including a State University of New York Chancellor's Award (1979), a National Catalyst Award for Excellence in Teaching (1992), the Estee Lectureship in Chemical

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