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1. In covalent bonds, electron sharing usually occurs so that atoms attain the electron configuration of noble gases. (8) 2. Atoms form double or triple covalent bonds if they can attain a noble gas structure by sharing two pairs or three pairs of electrons.

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8: Covalent Bonding and Molecular Structure

covalent bonds. The majority of covalent bonds form between nonmetallic elements. 8. Describe how the octet rule applies to covalent bonds. Atoms share valence electrons; the shared electrons complete the octet of each atom. 9. Illustrate the formation of single, double, and triple covalent bonds using Lewis structures.

Covalent Bonding Covalent Bonding

Covalent Bonding • There are several electrostatic interactions in these bonds: – Attractions between electrons and nuclei – Repulsions between electrons – Repulsions between nuclei • Covalent bond, sharing electrons, • But electron sharing not always equal.

Chapter 8 Concepts of Chemical Bonding - Chemistry

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This video describes how atoms covalently bond, and form single, double or triple bonds. Pi bonds are discussed as well as bond strength.

Chapter 8 Covalent Bonding Pt 1

Valence bond theory describes a covalent bond as the overlap of half-filled atomic orbitals (each containing a single electron) that yield a pair of electrons shared between the two bonded atoms. We say that orbitals on two different atoms overlap when a portion of one orbital and a portion of a second orbital occupy the same region of space.

8.1 Valence Bond Theory | Chemistry

A covalent bond is formed between atoms held together by sharing electrons. A molecule is a group of atoms joined by covalent bonds. Section 8.2 Assessment. PDF Chapter 8: Covalent Bonding. 242 Chapter 8 • Covalent Bonding Single Covalent Bonds When only one pair of electrons is shared, such as in a hydrogen molecule, it is a single

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