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$P_{\text{vap}} = \exp\left(\frac{14.790}{T} - \frac{5783}{T^2}\right)$  and by the ideal gas law  $f$  Solutions to Chemical and Engineering Thermodynamics, 3e  $n = \frac{P_{\text{vap}} V}{RT} = \frac{10^5 \text{ (bar)} \times 10^{-3} \text{ (m}^3\text{)}}{8.314 \text{ (J mol}^{-1}\text{K}^{-1}) \times 106 \text{ (K)}}$  volume taken up by liquid Equating  $P$  and  $P_{\text{vap}}$  we have  $x = \frac{10^5}{5432.8 \times 10^3}$ .

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Chapter 1 Solutions. Engineering and Chemical Thermodynamics. Wyatt Tenhaeff Milo Koretsky. Department of Chemical Engineering Oregon State University. [koretsm@engr.orst.edu](mailto:koretsm@engr.orst.edu) 1.2 An approximate solution can be found if we combine Equations 1.4 and 1.5:  $\frac{1}{2} m v_{\text{rms}}^2 = \frac{3}{2} kT$   $v_{\text{rms}} = \sqrt{\frac{3kT}{m}}$ . Assume the temperature is 22 C.

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