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Balancing Chemical Equations - Guided Inquiry Activity. Description. The learning goals of this guided-inquiry activity are as follows: Students will be able to: (1) Determine required conditions for a reaction to be considered “ balanced ” and relate these conditions to laws of matter, and (2) Develop strategies to balance chemical equations.

Balancing Chemical Equations - Guided Inquiry Activity ...

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BALANCING CHEMICAL EQUATIONS 1. 4. For each balanced reaction, indicate the total number of atoms in the table below.  
Reaction Total Number of Atoms Reactant Side (Left) Product Side (Right) Make Ammonia 8 8 Separate Water 6 6 Combust Methane 9 9  
5. Is the number of total atoms on the left side of a balanced equation always equal to the number of total atoms on the right side of the equation?

Balancing Chemical Equations Guided Inquiry\_StudentHandout ...

BALANCING CHEMICAL EQUATIONS 1 8. Start level 2 of the balancing equation game. Take turns in your group to balance the equations in the sim, using your strategies from Level 1, and adding new strategies as needed. Each person should be in charge of balancing at least one equation, asking for help from the group as needed.

Presley Mikulencak - Balancing Chemical Equations Guided ...

Explore the Balancing Chemical Equationssimulation. Discuss with your group what you find. a) What are the different ways that the simulation indicates when an equation is balanced? It shows the balance and then shows the amount of each element and when they are balanced it shows a happy face.

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The number of atoms of each element is the same on the left and right side of the equation because that is the goal of balancing equations. law of conservation of matter 7. As a group, play level 1 of the balancing equation game. Write down the strategies your group uses to balance chemical equations. BALANCING CHEMICAL EQUATIONS 1

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A balanced chemical equation represents a chemical reaction that obeys the Law of Conservation of Mass (Matter). It can be thought of as describing how many atoms or molecules of reactants are consumed in order to produce a certain number of atoms or molecules of products.

Balancing Chemical Equations

Balancing Chemical Equations - Inquiry based Introduction: Trish Loeblein: UG-Intro HS MS: MC Remote HW Lab: Chemistry:  
Balancing Chemical Equations - Guided Inquiry Activity: Timothy Herzog, Yuen-ying Carpenter: UG-Intro: Guided Remote HW:  
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### Balancing Chemical Equations - Chemical Equations ...

To balance a chemical equation, enter an equation of a chemical reaction and press the Balance button. The balanced equation will appear above. Use uppercase for the first character in the element and lowercase for the second character. Examples: Fe, Au, Co, Br, C, O, N, F.

### Chemical Equation Balancer

Balancing Chemical Equations - Guided Inquiry Activity. . The learning goals of this guided-inquiry activity are as follows: Students will be able to: (1) Determine required conditions for a reaction to be considered “ balanced ” and relate these conditions to laws of matter, and (2) Develop strategies to balance chemical equations.

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### Balancing Chemical Equations - Guided Inquiry Activity ...

As a group, play level 1 of the balancing equation game. Write down the strategies your group uses to balance chemical equations. 8. Start level 2 of the balancing equation game. Take turns in your group to balance the equations in the sim, using your strategies from Level 1, and adding new strategies as needed.

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BALANCING CHEMICAL EQUATIONS 2 8. Start level 2 of the balancing equation game and adding new strategies as needed.  
9. In the simulation, were you able to use noninteger numbers (like  $\frac{1}{2}$  or 0.43) for the coefficients

### BALANCING CHEMICAL EQUATIONS

The chemical equation has the products on the right side, while the reactants are written on the left side. Both of them are separated by an arrow. For instance,  $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$  denotes that there are four atoms of hydrogen and 2 atoms of oxygen on both sides of the equation. The amount of reactants must be equal to the amount of products.

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