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Arrange the following oxides in order of increasing acidity. Rank from least acidic to most acidic. To rank items as equivalent,

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overlap them.  $\text{CaO}$ ,  $\text{Al}_2\text{O}_3$ ,  $\text{SiO}_2$ ,  $\text{CO}_2$ ,  $\text{P}_2\text{O}_5$ ,  $\text{SO}_3$ . Part A barium oxide with water Express your answer as a chemical equation. Identify all of the phases in your answer.

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Rank from most to fewest valence  
electrons of the following  
elements: Kr, Ga, Br, As, K. -

Answer: Kr (8), Br (7), As (5), Ga

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(3), K (1). Valence electrons are equal to the atom's main group number. Link: <https://www.coursehero.com/file/20333255/Mastering-Chemistry-Chapter-2/>.

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