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and Rate of Decay

The halflife of a

particular nuclide is

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the time it takes
for half the nuclei
in a given sample
to decay This is
related to the

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Calculations:

Radioactive Decay

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How to calculate

how much of a

substance remains

after a certain

amount of time.

ALSO: How to

figure out how

Practice Problem:

Radioactive Half-

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Solution: Half life of
radioactive matters
depends on types
of nucleus or
neutron/proton
ratio. Physical
properties like
amount of matter,
temperature or
phase do not affect

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half life. Isotopes of same atom can have different half life since there are changes in their nucleus and neutron/proton ratio. II is true and I, III and IV are false. 5.

Nuclear Chemistry
(Radioactivity)
Exam2 and

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Problem Solutions
Half-Life Decay of a
Radioactive
Isotope; Half-Life
Percent of
Radioactive Isotope
Remaining; 0:
100.00: 1: 50.00:
2: 25.00: 3: 12.50:
4: 6.25: 5: 3.12: 6:
1.56: 7: 0.78: 8:
0.39: 9: 0.19: 10:
0.09

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Nuclear Chemistry:
Half-Lives and
Radioactive Dating

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y) How do you do
half life
calculations for
nuclear decay?
We'll do a whole
bunch of pra...

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Nuclear Half Life: Calculations - YouTube

Solution: $24.0 \text{ hr} /$
 $23.9 \text{ hr/half-life} =$
 $1.0042 \text{ half-lives}$

One day = one half-
life; $(1/2)^{1.0042} =$
 0.4985465

remaining = 4.98 g

Two days = two
half-lives; $(1/2)^{2.0084} =$

0.2485486

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remaining = 2.48 g

Seven days = 7
half-lives; $(1/2)^7$

$7.0294 =$

0.0076549

remaining =

0.0765 g

ChemTeam: Half-
Life Problems #1 -
10

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we're definite that

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Chemistry Half Life

Solutions Write the

nuclear equation

that represents the

radioactive decay

of radon-222 by

alpha particle

emission and

identify the

daughter isotope.

Solution. Radon

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has an atomic number of 86, so the parent isotope is represented as $\text{Th } 86 \text{ } 222 \text{ Rn}$.

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In this case we do not have an exact number of half-

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lives, so we need to use the more complicated equation (in Chapter 7 "Nuclear Chemistry", Section 7.2 "Half-Life") and solve for time. If the initial amount is represented by 16.0 mCi and the final amount is 5.6 mCi, we

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Half time is time required for half of mass of radioactive matter to decay. It is depends on types of matter or n/p ratio. If initial mass of matter is m_0 , after t time it has mass m , and if half life of matter is $t_{(1/2)}$; when $t=t_{(1/2)}$ $m=m_0/2$.

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Picture given below shows amount of mass as the time passes;

Half Life and
Radioactive Decay
Rates | Online
Chemistry ...

The half-life of
Carbon-14 is 5700
years. Calculate
the age of the
wood, in years. b)

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Suggest a reason why Carbon-14 is unsuitable for dating samples that are more than 50 000 years old.

15. 75% of the Potassium-40 atoms originally present in a rock sample were found to have undergone radioactive decay. The half-life of

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Potassium-40 is
 1.26×10^9 years.

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