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Thermodynamics: Conservation of Energy; Properties of Pure Substances (3 of 25)

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by Daniel V. Schroeder. Introduction This is an excellent textbook that is very easy to read and follow. The problems are enjoyable and interesting. I would recommend it highly. My Original Handwritten Notes. Download Problem Solutions (Part 1 of 5). Download Problem Solutions (Part 2 of 5). Download Problem Solutions (Part 3 of 5).

Introduction to Thermal Physics by Daniel V. Schroeder

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[Although Schroeder says we're not supposed to use the word "temperature" in the explanation, clearly what we're doing is decreasing the temperature of B and increasing that of A until the temperatures are equal. The slope  $\partial S = \partial U$  is just another way of referring to the temperature.]

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Schroeder, Introduction to Thermal Physics, An | Pearson

Thermodynamic identity (Schroeder 3.4)  $dU = TdS - pdV$  (1.5) Total differentials When we have a total differential, the things in front of the  $dS, dV$ , etc are partial derivatives.  $T = \left(\frac{\partial U}{\partial S}\right)_V$   $p = -\left(\frac{\partial U}{\partial V}\right)_S$  (1.6) Also, you can integrate along a path using a total differential, and can do linear algebra using total differential equations, e.g. substituting one for another.

Thermal Physics - Oregon State University

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Schroeder An Introduction To Thermal Physics

Zero-point energy (ZPE) is the lowest possible energy that a quantum mechanical system may have. Unlike in classical mechanics, quantum systems constantly fluctuate in their lowest energy state as described by the Heisenberg uncertainty principle. As well as atoms and molecules, the empty space of the vacuum has these properties. According to quantum field theory, the universe can be thought ...

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