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30.0 min 32.5 °C

27.5 °C 35 °C 45 °C.

45.0 min 36.2 °C

28.8 °C 40 °C 55 °C.

60.0 min 40 °C 30 °C

45 °C 65 °C. Step 1:

Create a line graph  
for each substance on  
graph below. Label

the substances. Step

2: Answer questions.

1. Order the

substances based. on

the time required to

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properly identify what kind of heating and cooling problem you are having, you will need to refer to a Worksheet Answers to Heat and Cooling Problems. A particular heat worksheet answers a specific problem you have. In fact, there are many different types of sheets that



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the following:

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practice Calculating  
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questions Higher

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Calculations

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involving Specific

Heat 1. For  $q = m c$

$T$  : identify each

variables by name &

the units associated

with it.  $q$  = amount of

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heat (J)  $m =$  mass  
(grams)  $c =$  specific  
heat ( $\text{J/g } ^\circ\text{C}$ )  $T =$   
change in  
temperature ( $^\circ\text{C}$ ) 2.  
Heat is not the same  
as temperature, yet  
they are related.  
Explain how they  
differ from each  
other.

Worksheet-  
Calculations

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Heat 1. For  $q = m c$

$T$  : identify each

variables by name &  
the units associated

with it.  $q$  = amount of  
heat (J)  $m$  = mass

(grams)  $c$  = specific

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change in  
temperature ( ° C) 2.  
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as

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just stop as soon as

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f, ice $5 = 3.34 \times 10 \text{ J/kg}$

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L. v, water. =  $22.5 \times 10^5$  J/kg.

Latent heat and Specific heat capacity questions. 1. How much water at  $50^\circ\text{C}$  is needed to just melt 2.2 kg of ice at  $0^\circ\text{C}$ ? 2. How much water at  $32^\circ\text{C}$  is needed to just melt 1.5 kg of ice at  $-10^\circ\text{C}$ ?

Latent heat and



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Specific heat capacity  
questions.

This is a GCSE  
worksheet from  
Flipped Around  
Physics, on specific  
latent heat. It covers  
rearranging and  
using the equation  
for specific latent  
heat. Worksheet  
answers are available  
from [www.flippedaroundphysics.com](http://www.flippedaroundphysics.com). The

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Worksheet is

designed to be used in a flipped learning setting. Students can be directed to the website ( [www.flippedaroundphysics.com](http://www.flippedaroundphysics.com)) to watch quality video material on the topic.

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Transfer/ Specific  
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Worksheet Solving

For Heat ( $q$ ) 1. How many joules of heat are required to raise the temperature of 550 g of water from 12.0 °C to 18.0 °C? 2.

How much heat is lost when a 64 g piece of copper cools from 375 °C, to 26 °C? (The specific heat of

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copper is  $0.38452 \text{ J/g}$   
 $\times \text{ }^\circ\text{C}$ ). Place your  
answer in kJ.

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Substance latent heat  
of fusion  $\text{J g}^{-1} \text{ }^\circ\text{C}^{-1}$   
melting temp. But

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worksheet

Heat is not the same  
as temperature yet

they are related. Use

$q = m \cdot \Delta t \cdot c_p$  to solve

the following

problems. Pin On For

Teacher . For  $q = m \cdot c \cdot \Delta t$ .

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worksheet answers. 2

what mass of water



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can be heated from  
25.0 °C to 50.0 °C by the  
addition of 2825

J. Worksheet  
calculations involving  
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Understanding  
specific heat capacity.  
GCSE worksheet on  
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Students use their understanding of specific heat capacity to explain why you burn your mouth on cheese when eating pizza. This diagnostic task asks students to think about the differences between temperature, kinetic energy stores and thermal energy stores and I hope,

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helps students to understand what specific heat capacity actually feels like!

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component of the  
NEMA 14-50R Heat  
System. It is  
essentially a  
document that serves  
as an instruction  
manual for  
installation of the  
system. In the event  
of a power outage,  
this manual will  
guide the installer  
through each and  
every step of the

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You need to  
understand how to  
project cash flow.  
Whatever your  
company planning  
objectives, cash flow  
remains the resource  
in the company, and

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cash is the one most important small business function.

Version control is another significant issue with Excel.

Heat Transfer  
Worksheet Answers -  
Briefencounters

100.0mL of  $4.0\text{ }^{\circ}\text{C}$   
water is heated until  
its temperature is  $37\text{ }^{\circ}\text{C}$ . If the specific



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heat of water is  $4.18 \text{ J/g } ^\circ\text{C}$ , calculate the amount of heat energy needed to cause this rise in temperature.  $25.0\text{g}$  of mercury is heated from  $25 \text{ } ^\circ\text{C}$  to  $155 \text{ } ^\circ\text{C}$ , and absorbs  $455$  joules of heat in the process. Calculate the specific heat capacity of mercury.

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